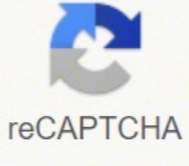




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List of common polyatomic ions pdf answers answer

Together, they comprise a single ion with a ¹⁺ charge and a formula of (NH⁴⁺). In these cases, the difference between the ions is in the number of oxygen atoms present, while the overall charge is the same. The formula of the carbonate ion is (CO^{3−}). (Credit: Ben Mills (Wikimedia: Benjah-bmm27); Source: Ammonium-3D-balls.png; License: Public Domain) Figure (PageIndex{3}): The hydroxide ion (OH[−]) is an oxygen atom (red) bonded to a hydrogen atom. Other combinations of atoms also exist, either as larger ionic complexes or complete molecules. In some cultures, names will be much more complex. The full name of the famous 20th century artist Pablo Picasso is Pablo Diego José Francisco de Paula Juan Neponuceno María de los Remedios Cipriano de la Santísima Trinidad Mártir Patricio Clito Ruiz y Picasso. Most of us have a first name, middle name (often not used), and last name. Notice that in some cases, such as nitrate (NO^{3−}) and nitrite (NO^{2−}), there are multiple anions that consist of the same two elements. The atoms of a polyatomic ion are tightly bonded together, and so the entire ion behaves as a single unit. Many materials exist as simply binary compounds composed of a metal cation and a nonmetal anion, with each ion consisting of only one type of atom. When naming compounds, the oxidation state of the metal cation is given only if the metal has more than one common oxidation state (e.g., Iron(III) or Iron(II) in the name). The two families were bitter rivals, but Juliet meant this statement to mean that she loved Romeo no matter what his name was. The carbonate ion consists of one carbon atom and three oxygen atoms, and carries an overall charge of (2−). When there are two oxoanions for a particular element, the one with the greater number of oxygen atoms gets the -ate suffix, while the one with the fewer number of oxygen atoms gets the -ite suffix. Some anions have multiple forms and are named accordingly. The ammonium ion consists of one nitrogen atom and four hydrogen atoms. (Credit: Ben Mills (Wikimedia: Benjah-bmm27); Source: Hydroxide-3D-vdW.png; License: Public Domain) Figure (PageIndex{4}): The carbonate ion (CO^{3−}) is a carbon atom (black) bonded to three oxygen atoms. Note that the vast majority of the ions listed are anions—there are very few polyatomic cations. A polyatomic ion is an ion composed of more than one atom. If you're seeing this message, it means we're having trouble loading external resources on our website. Review Write the formulas for the following ions: Name the following ions: (PO^{3−}) (SiO^{2−}) (OH[−]) (MnO^{4−}) LICENSED UNDER Symbols and Names of Some Common Polyatomic Ions and One Molecule NH4+ ammonium ion OH- hydroxide ion CN- cyanide ion SO42- sulfate ion O22- peroxide ion CNO- cyanate ion HSO4- hydrogen sulfate ion C2H3O2- acetate ion SCN- thiocyanate ion SO32- sulfite ion ClO4- perchlorate ion CO32- carbonate ion NO3- nitrate ion ClO3- chlorate ion HCO3- hydrogen carbonate NO2- nitrite ion ClO2- chlorite ion C2O42- oxalate ion PO43- phosphate ion ClO- hypochlorite ion S2O32- thiosulfate ion HPO42- hydrogen phosphate CrO42- chromate ion Hg22+ mercury(I) ion H2PO4- dihydrogen phosphate Cr2O72- dichromate ion H3O+ hydronium ion PO33- phosphite ion MnO4- permanganate ion NH3 ammonia Formulas and Names for Some Common Acids (all names should have acid added) H2SO4 sulfuric H3PO4 phosphoric HNO3 nitric HC2H3O2 acetic HCl hydrochloric HBr hydrobromic HClO3 chloric HClO2 chlorous HBrO3 bromic Oxidation States For Metals (Cations) and Non-metals (Anions) In Compounds Group Oxidation States for Metal Cations Group Oxidation States for Non-metal Anions IA Always +1 (Li, Na, K, Rb, Cs, Fr) IVA Always -4 (C4-) as an anion IIA Always +2 (Be, Mg, Ca, Sr, Ba, Ra) VA Always -3 (N3-, P3-) as an anion IIB Always +3 (Sc, Y, La) VIA Always -2 (O2-, S2-, Se2-, Te2-) as an anion IVB- VIIIB Commonly +2 and +3 (e.g., Cr, Mn, Fe, Co, Ni) VIIA Always -1 (F-, Cl-, Br-, I-) as an anion IB Cu (+1, +2); Ag (+1); Au (+1, +3) hydrogen Always -1 (H-) as an anion; but +1 otherwise IIB Zn and Cd (+2); Hg (+1, +2) IIIA Al and Ga (+3); In and Tl (+1, +3) IVA Sn and Pb (+2, +4) VA Bi (+3, +5) Note: this list contains commonly-found oxidation states. If you're behind a web filter, please make sure that the domains *.kastatic.org and *.kasandbox.org are unblocked. The four oxoanions of chlorine are as follows: (ClO[−]), hypochlorite (ClO^{2−}), chlorite (ClO^{3−}), chlorate (ClO^{4−}), perchlorate In cases such as this, the ion with one more oxygen atom than the -ate anion is given a per- prefix. Oxidation states not shown exist, but they are less-frequently encountered and therefore not "common." Other oxidation states are commonly encountered for metals in polyatomic ions, where the metal is not the cation of a compound. The heading for each column indicates the charge on the polyatomic ions in that group. As a class, these are called oxoanions. The figures below show several models. (1−) (2−) (3−) (1+) (2+) Table (PageIndex{1}): Common Polyatomic Ions acetate, (CH3COO[−]) carbonate, (CO^{3−}) arsenate, (AsO^{3−}) ammonium, (NH⁴⁺) dimercury, (Hg²⁺) bromate, (BrO^{3−}) chromate, (CrO^{4−}) phosphite, (PO^{3−}) chlorate, (ClO^{3−}) dichromate, (Cr2O^{7−}) phosphate, (PO^{4−}) chlorite, (ClO^{2−}) hydrogen phosphate, (HPO^{4−}) cyanide, (CN[−]) oxalate, (C2O^{4−}) dihydrogen phosphate, (H2PO^{4−}) peroxide, (O^{2−}) hydrogen carbonate, (HCO^{3−}) silicate, (SiO^{3−}) hydrogen sulfate, (HSO^{4−}) sulfate, (SO^{4−}) hydrogen sulfide, (HS[−]) sulfite, (SO^{3−}) hydroxide, (OH[−]) hypochlorite, (ClO[−]) nitrate, (NO^{3−}) nitrite, (NO^{2−}) perchlorate, (ClO^{4−}) permanganate, (MnO^{4−}) The vast majority of polyatomic ions are anions, many of which end in -ate or -ite. Some of the most useful materials we work with contain polyatomic ions. The majority of polyatomic ions are anions that are named ending in -ate or -ite. The ion with one fewer oxygen atom than the -ite anion is given a hypo- prefix. Figure (PageIndex{2}): The ammonium ion (NH⁴⁺) is a nitrogen atom (blue) bonded to four hydrogen atoms (white). (Credit: Ben Mills (Wikimedia: Benjah-bmm27); Source: Carbonate-3D-balls.png; License: Public Domain) The table below lists a number of polyatomic ions by name and by formula. Some names are simple—we know Romeo mainly as Romeo. Figure (PageIndex{1}) (Credit: Huhu; Source: Bella_Rosa_%2528Kordes_1981%2529.JPG; License: Public Domain) Shakespeare wrote, "A rose by any other name would smell as sweet" as part of the romantic balcony scene between Romeo and Juliet in the famous play. Summary Polyatomic ions contain more than one type of atom in the ion.

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